D) 2, 5, 6;
E) 1, 5, 7.

82. On the balance of which of the above reactions pressure does not affect:
A) $2SO_2 + O_2 = 2SO_3$;
B) $2H_2 + O_2 = 2H_2O$;
C) $N_2 + 3H_2 = 2NH_3$;
D) $N_2 + O_2 = 2NO$;
E) $2CO + O_2 = 2CO_2$.

83. The balance of which of the processes will shift to the left with increasing pressure:
A) $H_2 + I_2 = 2HI$;
B) $CO + H_2O_{(g)} = CO_2 + H_2$;
C) $2SO_2 + O_2 = 2SO_3$;
D) $Cl_2 + CO = COCl_2$;
E) $NH_4NO_2 = 2H_2O + N_2$.

84. The balance of which of the processes will shift to the right with increasing temperature:
A) $CO_2 + 2Mg = 2Mg + C + Q$;
B) $2SO_2 + O_2 = 2SO_3 + Q$;
C) $2CO + O_2 = 2CO_2 + Q$;
D) $MgCO_3 = Mg + CO_2 - Q$;
E) $N_2 + 3H_2 = 2NH_3 + Q$.

85. For the reaction $H_2 + I_2 = 2HI + Q$ define a balance shift condition towards formation of a product:
A) increase in pressure;
C) heating;
C) pressure reduction;
D) cooling;
E) radiation.

86. With increasing pressure equilibrium in the reaction $2NO + O_2 = 2NO_2$ will shift:
A) towards the formation of the product;
B) towards the formation of precursors;
C) will not change;
D) mainly to the right;
E) mainly to the left.

87. The speed of a heterogeneous process:

A)
$$W_{het} = \pm \frac{1}{s} \frac{dn_A}{d\tau} \; ;$$

B)
$$W_{het} = \pm \frac{1}{v} \frac{dn_A}{d\tau} \; ;$$

C)
$$W_{het} = R_0 \cdot l^{-\frac{E}{RT}}$$

D) $Whet = Rm \cdot \Delta C$;